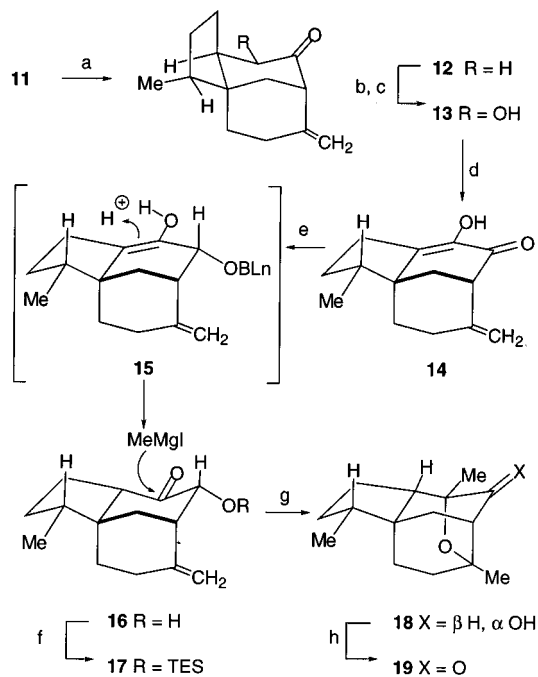




## Scheme 2



<sup>a</sup> Reagents: (a) Li/NH<sub>3</sub> (85%); (b) triethylsilyl triflate (TESOTf), NEt<sub>3</sub>; (c) *m*-chloroperoxybenzoic acid (*m*-CPBA), NaHCO<sub>3</sub> (79% from **12**); (d) trifluoroacetic anhydride (TFAA), DMSO, NEt<sub>3</sub> (80%); (e) NaBH<sub>4</sub>, 30 s (84% based on recovered **14**); (f) TESOTf, NEt<sub>3</sub> (93%); (g) MeMgI, then 1.0 N HCl in Et<sub>2</sub>O, then 5 N aqueous HCl (56%); (h) Jones reagent (85%).

action of mercuric chloride, as shown.<sup>12,13</sup> Reductive cleavage of the vinyl mercurial and deprotection of the vinylsilane were accomplished in the manner indicated to afford **11** (*cf.* hypothetical construct **3** where Y = H and X = O).

Birch-like reduction of the enone linkage afforded the cis-fused dihydro product **12** (Scheme 2). This result presumably reflects substantial early rehybridization at C9 in the direction of a conformer in which the emerging hydrindanone is pre-cisoid.<sup>14</sup> In general, metal–ammonia reductions of non-octalone ring systems tend to afford the thermodynamic product.<sup>15</sup> Hydroxylation at C10 was accomplished via Rubottom<sup>16</sup> oxidation of the silyl enol ether derived from **12** (see compound **13**).

Oxidation<sup>17</sup> of the C10 alcohol of **13** resulted in the formation of diosphenol **14** (*cf.* **3** X = O; Y = OH). Fortunately, the required  $\beta$ -stereochemistry at C9 could be constituted by treatment of this compound with sodium borohydride. Reduction occurred by hydride delivery from the  $\beta$ -face (see proposed intermediate **15**). *Most importantly, disruption of the diosphenol connectivity triggered ketonization of the C9–C10 enol with protonation at C9 occurring strictly from the  $\beta$ -face (see compound **16**).* We note that this sense of protonation is in contrast to the  $\alpha$ -mode encountered in the Birch reduction step (*cf.* **11**  $\rightarrow$  **12**). For the enol linkage of **15**, the sense of protonation is governed by steric accessibility factors that would certainly favor reactions from the exo face of the bicyclo[3.3.1]-moiety.

At this point, the C11 hydroxyl group of **16** was protected in the form of its triethylsilyl ether (see compound **17**). Not unexpectedly, reaction of the keto function of **17** with methylmagnesium iodide resulted in introduction of the methyl group from the  $\beta$ -face and appearance of an  $\alpha$ -disposed hydroxyl

(12) Drouin, J. D.; Boaventura, M.-A.; Conia, J.-M. *J. Am. Chem. Soc.* **1985**, *107*, 1726.

(13) (a) Huang, H.; Forsyth, C. J. *J. Org. Chem.* **1995**, *60*, 2773. (b) Huang, H.; Forsyth, C. J. *J. Org. Chem.* **1995**, *60*, 5746.

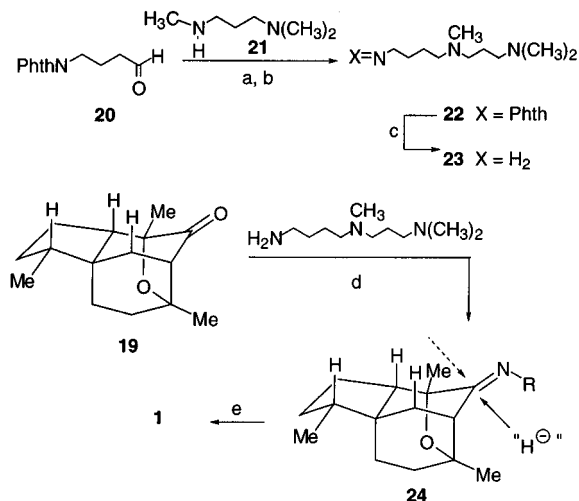
(14) Stork, G.; Darling, S. D. *J. Am. Chem. Soc.* **1960**, *82*, 1513.

(15) Caine, D. *Org. React. (N.Y.)* **1976**, *23*, 1.

(16) Rubottom, G. M.; Gruber, J. M. *J. Org. Chem.* **1978**, *43*, 1599.

(17) Kawada, K.; Gross, R. S.; Watt, D. S. *Synth. Commun.* **1989**, *5&6*, 777.

## Scheme 3



<sup>a</sup> Reagents: (a) **21** (1.1 equiv), benzene, reflux, (Dean–Stark); (b) NaCNBH<sub>3</sub>, 1:1 THF/MeOH, pH 4 (50% from **20**); (c) MeOH/NH<sub>2</sub>NH<sub>2</sub>, 4:1 (55%); (d) **23** (6.0 equiv), toluene, pyridinium *p*-toluenesulfonate (PPTS) (cat.), powdered sieves, reflux, 3 days; (e) NaCNBH<sub>3</sub>, MeOH, pH 4 (82% from **19**).

group. Treatment of the crude Grignard product with ethereal HCl brought about formation of the caged ether structure. Subsequent addition of aqueous HCl effected desilylation (see compound **18**). We attribute the particularly facile cyclic ether formation to the enforced proximity of the tertiary hydroxyl disposed on the endo face of the bridged system with the exocyclic methylene at C12. Jones oxidation<sup>18</sup> of **18** produced ketone **19**.

Aldehyde **20**<sup>19</sup> was utilized to reductively alkylate commercially available diamine **21** under the protocols described by Borch (Scheme 3).<sup>20</sup> Dephtaloylation of the resultant **22** afforded **23**. Condensation of ketone **19** with triamine **23**, as shown, led to **24**. Finally, reduction of **24** with sodium cyanoborohydride afforded (*dl*)-hispidospermidin (**1**). It will be noted that in this reduction, the hydride equivalent was delivered from the  $\alpha$ -face of C11. This result need occasion little surprise since attack at the C11 imine is occurring through a trajectory which does not involve incursion into the concave face of the cage-like domain. Unlike the hypothetical trajectory associated with axial delivery of hydride to C11 (which must pass through potentially deflecting axial hydrogens at C9 and C13, as well as the equatorial  $\beta$ -methyl at C10), the pathway for equatorial face reduction of the imine is relatively unencumbered. Furthermore, the bridging oxygen may well be exerting a directive effect.

The structure of the hispidospermidin generated by total synthesis follows from its 400 MHz NMR spectrum in conjunction with the high-field spectra of its precursors. Moreover, the 400 MHz NMR spectrum of an authentic sample of hispidospermidin was identical with that of synthetic racemate. Clearly, ketone **19** provides easy access for introduction of a variety of basic side chains at C11 using the stereochemical model established in the total synthesis.

**Acknowledgment.** We dedicate this paper to the memory of Paul Dowd. This work was supported by the National Institutes of Health (HL 25848). We are grateful to Vinka Parmakovich and Barbara Sporer of the Columbia Mass Spectral Facility. We also thank the Nippon Roche Corp. for providing us with a sample of the natural product.

**Supporting Information Available:** Selected experimentals (for compounds **11**, **16**, and **18**) and structural information (for compounds **13** and **16**) (4 pages). See any current masthead page for ordering and Internet access instructions.

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(18) Meinwald, J.; Crandall, J.; Hymans, W. E. In *Organic Syntheses*; New York, 1973; Collect. Vol. V, p 866.

(19) Hamilton, R.; Walker, B. J.; Walker, B. *Tetrahedron Lett.* **1993**, *34*, 2847.

(20) Borch, R. F.; Hassid, A. I. *J. Org. Chem.* **1972**, *37*, 1673.